

Closed-Loop Polymer Upcycling by Installing Property-Enhancing Comonomer Sequences and Recyclability

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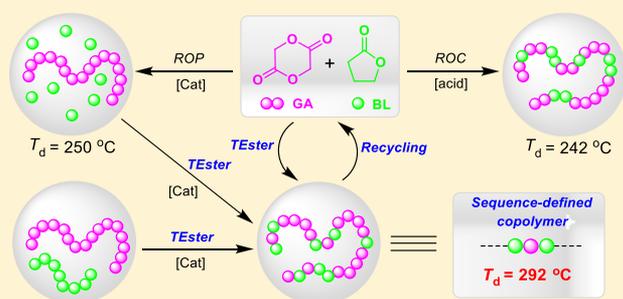
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Supporting Information

ABSTRACT: The concept of upcycling postconsumer plastics into higher-value products is attractive, but the challenges remain to develop a cost-effective upcycling scheme, discover property-enhancing structures, and, most importantly, install recyclability into upcycled plastics to enable a circular lifecycle. Reported herein is a convenient and effective strategy to upcycle polyester, exemplified by poly(glycolic acid) (PGA), via transesterification (TEster) in bioderived, commercially available γ -butyrolactone (BL) that serves as both the solvent and comonomer, which generates sequence-defined copolymer poly(GA-co-BL). Owing to the isolated glycolic sequence present in the copolymer created uniquely by TEster, it exhibits much-enhanced thermal stability (≥ 44 °C) over both homopolymers or copolymers without such sequences. This upconverted copolymer is chemically recyclable, enabling a complete recovery of pure glycolic acid and BL feedstocks.



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INTRODUCTION

Current methods for recycling of postconsumer plastics are largely ineffective due to minimally retained material's values, thus presenting a pressing need and also an innovation opportunity for developing new approaches to recycle plastics more effectively, en route to an end goal of achieving a circular materials economy.^{1–21} In this context, the concept of upcycling is an attractive approach, as, in principle, higher-value or value-added new materials would be the end products of the upcycling or repurposing process.^{22–30} However, the challenges have been either the need to use expensive reagents or fine chemicals and/or elaborated conditions for the upconversion or postmodification of the existing polymers into the higher-value products or the generation of new materials waste when the upcycled polymers are no more recyclable than the previous polymers.^{22–30}

Copolymerizing two or more monomers is a well-established strategy to create new copolymer materials that often deliver unattainable or improved properties over those of homopolymers.^{31–34} By adjusting the comonomer chemical nature, sequence, stereochemistry, and the composition, the properties and functions of the resultant copolymers can be tailored in a wide range to meet the desired application demand.^{35–45} A notable example in this context is poly(lactide-co-glycolide) (PLGA),^{46–54} which was approved by FDA for biomedical applications and shows improved flexibility, processability, and degradation behavior, relative to homopolymer poly(glycolic

acid) (PGA), since PGA's melting-transition temperature (T_m) of ~ 219 °C is close to its onset thermal decomposition temperature (T_d) of ~ 250 °C, and its dissolvability in common organic solvents is extremely poor.^{46,47} Nevertheless, the resulting PLGA is still a rather brittle material ($\sim 6\%$ elongation at break).

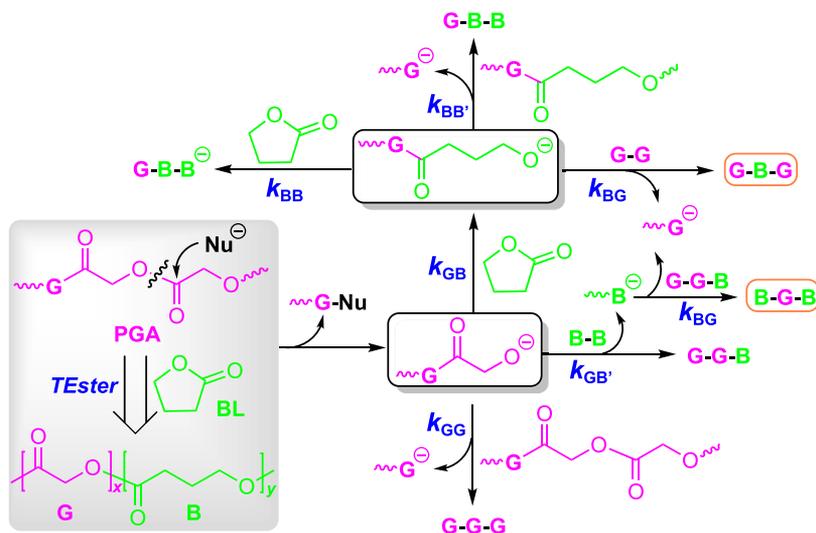
By exploiting the unique ability of the copolymerization to create new or better materials, we reasoned that plastic upcycling could be achieved through the formation of better-performing copolymers that incorporate property-enhancing functions or structures into the resulting copolymer. However, to realize a closed-loop upcycling scheme, three challenges must be addressed: (a) to develop a synthetic pathway that can produce such copolymers in a convenient and cost-effective manner; (b) to uncover property-enhancing copolymer structures or functions; and (c) to install the recyclability into the upcycled copolymers. Here, we present a proof-of-concept example to demonstrate the feasibility of such an upcycling strategy. In this example of upcycling of polyester, represented by PGA, it was simply treated with bioderived γ -butyrolactone (BL) as both solvent and comonomer to produce sequence-defined copolymer poly(GA-co-BL) with BL incorporation up to 20.0% (the percent BL relative to the

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Scheme 1. Proposed TEster Pathways to Produce Poly(GA-co-BL) with Various Possible Sequences (G = 1/2GA, B = BL Ring-Opened Unit) Starting from Homopolymer PGA and Comonomer BL, where k_{GG} Defines the Rate Constant for the G-Centered Anion-Attacking G-Chain Carbonyl, whereas k_{GB} and $k_{GB'}$ Represent the Rate Constant for the G-Centered Anion-Attaching B (BL) Monomer and B' (B Chain Carbonyl), Respectively^a



^aLikewise, k_{BB} , $k_{BB'}$, and k_{BG} are defined similarly.

Table 1. Results of the TEster between PGA and BL^a

run	cat.	I.	temp (°C)	solvent	time (day)	yield ^b (mg)	incorp. ^c (BL %)	M_n _{NMR} ^d (kg/mol)	M_n _{GPC} ^e (kg/mol)	\bar{D} ^e M_w/M_n
1	La	BnOH	25	THF	7	95	5.00	3.70		
2	La		25	THF	7	98	1.00	<i>f</i>		
3	TBD		25	DCM	7	96	2.70	<i>f</i>		
4	TBD	BnOH	25	DCM	7	95	3.23	4.04		
5	DBU		25	DCM	7	97	1.41	<i>f</i>		
6	DBU	BnOH	25	DCM	7	96	2.86	4.27		
7	^t Bu-P ₄		25	THF	7	99	0.61	<i>f</i>		
8	^t Bu-P ₄	BnOH	25	THF	7	96	2.43	4.28		
9	La	BnOH	25	BL	7	98	5.43	3.85		
10	La	BnOH	80	BL	1	96	2.83	10.6		
11	La	BnOH	120	BL	1	100	8.34	6.20		
12	La	BnOH	160	BL	1	103	14.3	6.50	6.38	1.25
13	La	BnOH	160	BL	0.5	99	11.7	5.70	4.68	1.19
14	La	BnOH	160	BL	1.5	105	18.6	6.60	6.51	1.22
15	La	BnOH	160	BL	2	107	20.0	6.80	6.61	1.25
16	TBD	BnOH	160	BL	1 h	100	7.00	8.58		
17	TBD	BnOH	160	BL	4 h	102	16.8	7.60	7.25	1.20
18	TBD	BnOH	160	BL	6 h	106	17.2	4.23	6.50	1.15
19	La	BnOH	160	BL	2	104	19.3	6.92	6.80	1.30

^aConditions for runs 1–8: cat. = 51.6 μ mol, PGA = 100 mg, BL = 390 mg, [BL + PGA]/cat./I. = 100/1/2, BL/PGA = 5, solvent = 0.5 mL; conditions for runs 9–18: cat. = 5.16 μ mol, BnOH = 10.32 μ mol, PGA = 100 mg, BL = 1.0 mL; conditions for run 19: cat. = 5.16 μ mol, BnOH = 10.32 μ mol, PGA = 100 mg, PBL = 113 mg, BL = 0.9 mL. ^bIsolated yield. ^cBL incorporation of the copolymer was measured by ¹H NMR spectrum (BL mol % = $[I_B/(I_B + I_G)] \times 100\%$ = $[I_{2.14\text{ppm}}/(I_{2.14\text{ppm}} + I_{4.92-5.02\text{ppm}})] \times 100\%$, G = glycolyl (–COCH₂O–)). ^dCalculated by ¹H NMR spectra: M_n (g/mol) = $(58.03 \times I_{4.92-5.02\text{ppm}} + 86.09 \times I_{2.14\text{ppm}})/I_{5.31\text{ppm}} + 108.13$ (BnOH). ^e M_n and \bar{D} were determined by GPC at 40 °C in DMF/LiBr relative to poly(methyl methacrylate) standards. ^fThe signals of chain-end groups were difficult to detect in ¹H NMR spectrum.

glycolyl (COCH₂O) unit, abbreviated as G, which is 1/2 glycolide) via catalyzed transesterification (TEster). Interestingly, this copolymer exhibits much-enhanced thermal stability (≥ 44 °C) over both homopolymers (PGA and PBL) or copolymers without such sequences, thanks to the presence of isolated G sequences [G–B–G or B–G–B; B = BL ring-opened unit] unique to the copolymer produced by TEster. Also importantly, this upcycled poly(GA-co-BL) can be completely recycled back to its building blocks, glycolic acid

and BL, thus successfully constructing a closed-loop upcycling scheme.

RESULTS AND DISCUSSION

Upconversion of PGA to Copolymer by TEster. The choice of BL as both the comonomer and solvent for upcycling of PGA was primarily based on two reasons. First, it is a bioderived monomer or industrial solvent, readily available in large quantities,⁵⁵ and its good solubility for many chemicals

and polymers could potentially dissolve long crystalline GA blocks to ensure the formation of copolymers with a homogeneous composition. Second, it has been established that PBL, synthesized via ring-opening polymerization of nonstrained BL under suitable thermodynamic, kinetic, and processing conditions with metal-based or organic catalysts,^{15,16,56–60} can be selectively and quantitatively recycled back to BL by thermolysis or chemolysis.^{15,16} Therefore, the utilization of BL may enable or enhance the chemical recyclability of the upcycled copolymer.

At the outset, we needed to synthesize the target poly(GA-co-BL) copolymer to examine its properties and chemical recyclability and in turn answer the question of whether this material constitutes an upcycled product or not. Direct ring-opening copolymerization (ROC) of highly reactive GA with highly stable BL was reported to be successful only by cationic polymerization with acidic initiators, producing copolymers in low yields (e.g., 1.4% by FeCl₃ after 44 h at 60 °C; 26% by Sn(Ph)₄ after 96 h at 140 °C).^{61,62} Metal-mediated or organic-mediated anionic ROC of BL with moderately reactive lactones such as ϵ -caprolactone was successful.^{63,64} Consistent with the very large reactivity difference between GA and BL toward ROC, our attempted synthesis of poly(GA-co-BL) via ROC of GA and BL by either coordination or anionic polymerization with metal-based La[N(SiMe₃)₂]₃ (with or without BnOH) and organic phosphazene superbase tBu-P₄/BnOH catalyst/initiator systems was unsuccessful, producing exclusively PGA homopolymer with no BL incorporation, even with excess BL (BL/GA = 10/1) and at low temperature (−40 °C, the condition suitable for facile ROP of BL) (Figure S1). The use of other metal catalysts such as the discrete yttrium complex supported by the amino-alkoxy-bisphenolate ligand⁶⁵ as well as other commonly used organic catalysts [e.g., 1,5,7-triazabicyclo[4.4.0]dec-5-ene (TBD), 1,8-diazabicyclo[5.4.0]undec-7-ene (DBU)]^{66–68} gave the same result. By employing the strategy developed for the synthesis of random PLGA,⁵⁴ the copolymerization was also performed by slow and continuous addition of a GA solution to BL after initiating BL first; however, the resulting product turned out to be the mixture of PGA and PBL, rather than the anticipated copolymer (Figures S2 and S3). The use of BnOH as the initiator in the cationic copolymerization promoted by BF₃·Et₂O and CF₃SO₃H can improve the yield and BL incorporation (up to 33.8%, Table S1 and Figure S4), but the resulting copolymer is an inferior material with a noticeably lower T_d (241 °C) than PGA.

The above largely unsuccessful attempts pointed to a need to develop a new synthetic pathway by employing homopolymer PGA directly. In this context, we reasoned that the TEster between PGA and BL with a suitable catalyst could afford the copolymer according to the pathways outlined in Scheme 1. The TEster between PGA and BL was examined with different catalysts and solvents at 25 °C (runs 1–8, Table 1), successfully affording poly(GA-co-BL) with BL incorporations up to 5.0% by the best-performing catalyst/initiator La[N(SiMe₃)₂]₃/BnOH in tetrahydrofuran (THF) (run 1, Table 1), albeit for a long time (7 days). Next, we performed the TEster of PGA in BL that serves as both the solvent and comonomer at different temperatures, rendering the efficient synthesis of poly(GA-co-BL) with BL incorporations ranging from 2.83 to 20.0% while reducing the reaction time to 12 h when performed at 160 °C (runs 9–15, Table 1). The TEster by TBD/BnOH was even faster (runs 16–18, Table 1),

producing poly(GA-co-BL) with a high BL incorporation up to 17.2% in 6 h. The resultant copolymers with BL incorporations above 12.0% all exhibit good dissolvability in dimethylformamide (DMF), which enabled measurements of molecular weights by gel-permeation chromatography (GPC), showing a somewhat heterogeneous composition of the copolymer at the early stage of the TEster and then gradually becoming homogeneous at the late stages (Figure S5), as indicated by unimodal distribution with relatively low dispersity (e.g., M_n = 6.61 kg/mol, \bar{D} = 1.25 for the copolymer with 20.0% BL incorporation). However, for the copolymers with low BL incorporations (<12%), the dissolvability in common organic solvents (e.g., DMF, THF, CHCl₃, CH₂Cl₂) is extremely poor. Corrosive fluorinated solvents (e.g., trifluoroacetic acid) are required to dissolve these copolymers. Even in hexafluoroisopropanol, such copolymers are only partially soluble. Hence, the poor dissolvability of these copolymers makes GPC measurements infeasible. Noteworthy, here the TEster between two homopolymers PGA and PBL by La[N(SiMe₃)₂]₃/BnOH was also successful; monitoring the reaction by ¹H NMR showed that it started with fast, quantitative depolymerization of PBL into BL, followed by the TEster between PGA and BL in which the activity and BL incorporation (run 19, Table 1) were similar to the TEster starting with PGA and BL. To determine if the copolymer by the cationic ROC of GA and BL with BF₃·Et₂O/BnOH was also produced by the TEster mechanism, we performed the kinetic study. Specifically, extending the polymerization time from 2 to 12 h, the copolymer yield increased from 16.0 to 43.6%, while the BL incorporation was maintained at around 20% (runs 5, 8–10, Table S1), indicating the copolymer by cationic polymerization was not obtained by the TEster pathway.

Property-Enhancing Comonomer Sequences. ¹H and ¹³C NMR analyses of the selected poly(GA-co-BL) samples showed different copolymer microstructures produced by the TEster with La[N(SiMe₃)₂]₃/BnOH from those by cationic ROC with BF₃·Et₂O/BnOH.^{61,62} Because the chemical shifts of the glycolyl unit (G) and BL ring-opened unit (B) are sensitive to electronic differences, the detailed sequence information of triad glycolyl units and dyad BL ring-opened units can be revealed by ¹H NMR. Thus, for the copolymers formed by TEster, the peaks at δ 5.02, 4.92–4.95, and 4.86 ppm (Figure 1b,c) are attributed to the CH₂ protons of the glycolyl unit with a G–G–G homosequence, B–G–G/G–G–B heterosequence, and B–G–B alternating sequence, respectively. The signals of the CH₂ protons of the BL ring-opened unit split into two sets of peaks, in which the predominant peaks at δ 4.40, 2.72, and 2.17 ppm correspond to B–G alternating sequence while minor peaks at δ 4.32, 2.61, and 2.11 ppm are attributed to the B–B homosequence.

Table 2 lists the calculated molar fractions of different sequences based on ¹H NMR spectra. With increasing BL incorporation in the copolymer produced by the TEster with La[N(SiMe₃)₂]₃/BnOH from 5.00 to 20.0%, the molar fraction of the G–G–G homosequence decreased from 93.5 to 60.8%, while that of the isolated G in the form of B–G–B alternating sequence enhanced from 0.65 to 4.50%. In contrast, the content of the B–B homosequence was kept at the constantly low level of ~7.0%, independent of the BL incorporation (entries 1–4, Table 2). The copolymer obtained by the TEster with TBD/BnOH had similar sequence distributions (entry 5, Table 2). However, the microstructure of the copolymer

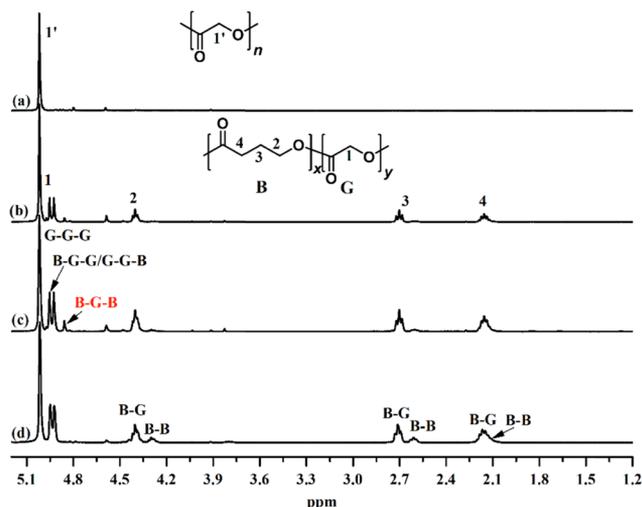


Figure 1. ^1H NMR spectra (CF_3COOD) of poly(GA-co-BL)s: (a) BL = 0% (PGA for comparison); (b) BL % = 14.3% (run 12, Table 1); (c) BL % = 20.0% (run 15, Table 1); and (d) BL % = 20.2% (produced by ROC, run 5, Table S1).

produced by the cationic ROC is *substantially different*, in which the B–G–B alternating sequence was not detected and the content of the B–B homosequence was much higher (23.0%, Figure 1d; entry 6, Table 2). Overall, the low content of the B–B homosequence and the presence of the appreciable amount of the unique B–G–B alternating (i.e., isolated G) sequence indicate that poly(GA-co-BL) formed by TEster has a much more sequence-defined structure than that formed by ROC. The observed low content of the B–B homosequence in the copolymer is presumably a result of a combination of two factors. First, the nucleophilic attack of the B-centered anion onto ring-opened B units ($k_{\text{BB}'}$ pathway, Scheme 1) is infrequent. Second, the nucleophilic attack of the B-centered anion onto the monomer B molecules (k_{BB} pathway, Scheme 1) is strongly competed by the backbiting ring-closed degradation reaction of $\cdots\text{B}-\text{B}-\text{B}-\text{OH}$ that has a low transition state (TS) energy, as indicated by our computational study (vide infra). The same sequence structures were also observed from the ^{13}C NMR spectra (Figure S6), thus providing corroborative evidence for the microstructure assignments.

The thermal properties of poly(GA-co-BL)s with different BL incorporations and sequences were examined by differential scanning calorimetry and thermal gravimetric analysis (TGA). As compared to PGA, which displays a T_m of 219 °C without a noticeable T_g due to its high crystallinity, both T_m value and

crystallinity of the copolymers decreased gradually with increasing BL incorporation, coupled with an appearance of T_g transition (Figures S7 and S8), highlighting the plasticizing effect of the incorporated BL units in the copolymer. Noteworthy also is the observation of a single T_m for the copolymer, further confirming the homogeneous composition of the copolymer. More significantly, poly(GA-co-BL)s produced by the TEster and ROC methods exhibited largely different thermostability, in reference to that of PGA (Figure 2). In the case of poly(GA-co-BL) formed by ROC ($M_{n\text{GPC}} =$

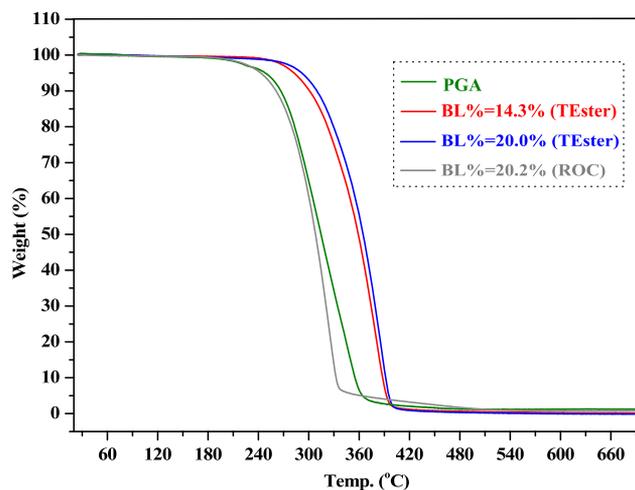


Figure 2. TGA curves: PGA (green); poly(GA-co-BL) with BL % = 14.3% produced by TEster (red); poly(GA-co-BL) with BL % = 20.0% produced by TEster (blue); poly(GA-co-BL) with BL % = 20.2% produced by ROC (gray).

6.80 kg/mol, 20.2% BL incorporation), the T_d value is 241 °C, which is somewhat lower than that of PGA ($T_d = 248$ °C). In sharp contrast, at the very similar level of molecular weight ($M_{n\text{GPC}} = 6.61$ kg/mol) and BL incorporation (20.0%), poly(GA-co-BL) produced by TEster has a high T_d value of 292 °C, which is 51 and 44 °C higher than that of the copolymer formed by ROC and PGA, respectively.

The different thermostability between the two types of copolymers can be attributed to different monomer sequence structures, but an important fundamental question is *how the copolymer sequences affect the thermostability*. In principle, poly(GA-co-BL) presents three types of cleavable sites during thermal degradation, including cleavages of continuous B–B and G–G homosequences and B–G alternating sequences. As the T_d value of PGA is higher than that of the linear PBL (208

Table 2. Molar Fractions of Different Sequence Structures Calculated by ^1H NMR Spectra

entry	incorp. (%)	glycolyl unit (%) ^d			BL-ring-opened unit (%) ^d	
		G–G–G	G–G–B/B–G–G	B–G–B	B–G	B–B
1 ^a	5.00	93.5	5.85	0.65	92.6	7.40
2 ^a	14.3	73.6	24.6	1.80	93.5	6.50
3 ^a	18.6	64.4	32.4	3.20	92.9	7.10
4 ^a	20.0	60.8	34.7	4.50	92.9	7.10
5 ^b	16.8	65.4	31.6	3.00	93.3	6.70
6 ^c	20.2	60.8	39.2	0	77.0	23.0

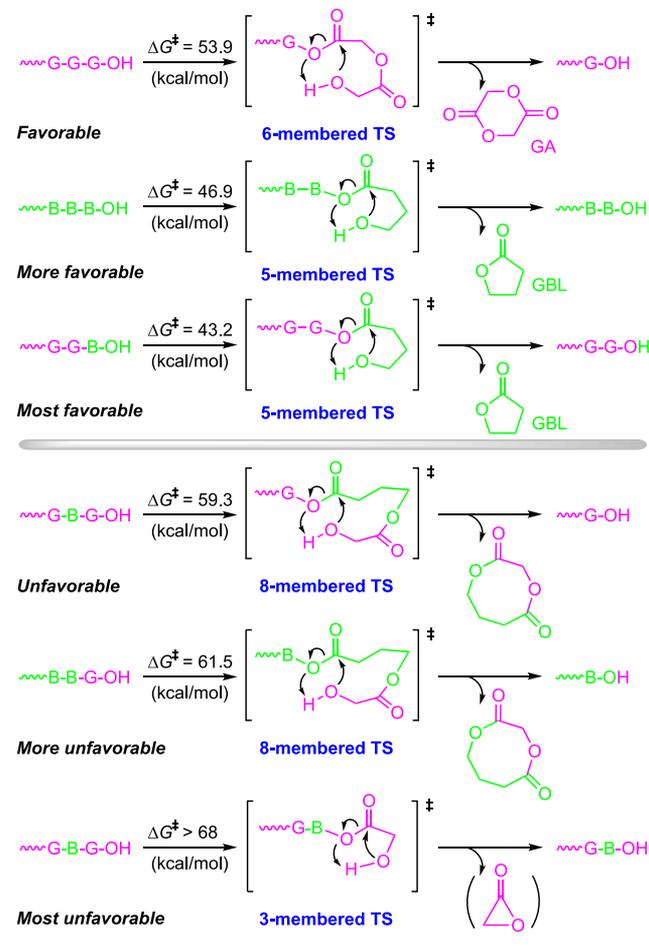
^aPoly(GA-co-BL)s produced by TEster with $[\text{La}]/\text{BnOH}$. ^bPoly(GA-co-BL) produced by TEster with TBD/BnOH . ^cPoly(GA-co-BL) produced by ROC with $\text{BF}_3\cdot\text{Et}_2\text{O}/\text{BnOH}$. ^dMolar ratios of different sequence linkages determined by ^1H NMR spectra (see Supporting Information (SI) for details).

°C) but considerably lower than that of the poly(GA-co-BL) formed by TEster that mainly consists of G-G continuous and B-G alternating sequences; it can be concluded that the thermal stability of the sequence structures follows the order of B-G > G-G > B-B. This simple analysis suggests that the more sequence-defined the structure the copolymer is, the higher the thermal stability it possesses. The copolymer formed by ROC contains a high content of B-B continuous sequence (23.0%) without any detectable isolated G sequence (entry 6, Table 2), while the copolymer formed by the TEster contains 4.5% of the B-G-B alternating sequence and a much-reduced amount of the B-B continuous sequence (7.1%, entry 4, Table 2). Therefore, the poly(GA-co-BL) formed by TEster exhibits a much higher thermal stability than the poly(GA-co-BL) formed by ROC, thanks to the generated isolated G sequence and the reduced B-B continuous sequence. Moreover, the T_d value of the TEster copolymer enhances with increasing BL incorporation (e.g., BL % = 14.3%, T_d = 280 °C; BL % = 20.0%, T_d = 292 °C), which can also be explained by a more sequence-defined structure as a result of an increase of the B-G-B sequence from 1.8 to 4.5%.

Computational Study of Sequence/Thermostability Relationship. We also addressed this sequence/thermostability relationship through density functional theory (DFT) calculations (see the SI for computational details) performed on the transition-state (TS) energies involved in the possible decomposition pathways of the chains with various sequences (Scheme 2). Focusing first on the chains with continuous G and B sequences carrying the -OH end group, represented by $\cdots\text{G-G-G-OH}$ and $\cdots\text{B-B-B-OH}$, the calculated ΔG^\ddagger of the 6- and 5-membered TSs accounting for the first degradation step of the PGA and PBL chains was 53.9 and 46.9 kcal/mol, respectively. This result suggests that PGA is more thermally resistant, thus possessing higher thermostability than PBL, consistent with the observed T_d values: 248 °C for PGA and 208 °C for PBL. Replacing the third and second G unit in $\cdots\text{G-G-G-OH}$ with a B unit gave $\cdots\text{B-G-G-OH}$ and $\cdots\text{G-B-G-OH}$ sequences, with the calculated ΔG^\ddagger values of 55.4 and 59.3 kcal/mol for their respective 6- and 5-membered degradation TS (Scheme S1). This result is intriguing, implying that the copolymer with isolated G units should be more thermally stable than both homopolymers, also consistent with the experimental observations (vide supra). Noteworthy, the sequence resulting from the replacement of both the second and third G units down the chain, $\cdots\text{B-B-G-OH}$, was computed to be even more resistant to degradation, with a higher ΔG^\ddagger value of 61.5 kcal/mol for the 8-membered TS; this is a result of the interplay between the conformation of the ring-unzipping TS and the strength of the breaking/forming bonds, as the ester bond involving a B unit is stronger than the ester bond involving a G unit (see the SI). An alternative degradation pathway for the isolated G unit located at the chain end is through a 3-membered TS, but the free energy required is the highest, with $\Delta G^\ddagger > 68$ kcal/mol (Scheme 2).

Likewise, substituting the B units in the starting chain $\cdots\text{B-B-B-OH}$ with a G unit gave chains with sequences of $\cdots\text{G-B-B-OH}$, $\cdots\text{B-G-B-OH}$, and $\cdots\text{G-G-B-OH}$. The effect of the third G unit on the decomposition TS energy is negligible, but the second G unit lowered the TS further by 3.7 kcal/mol (second vs third entry, Scheme 2) due to the breaking of the weaker G ester bond. As for the chain carrying a B-OH terminal group, elimination of a B unit is relatively

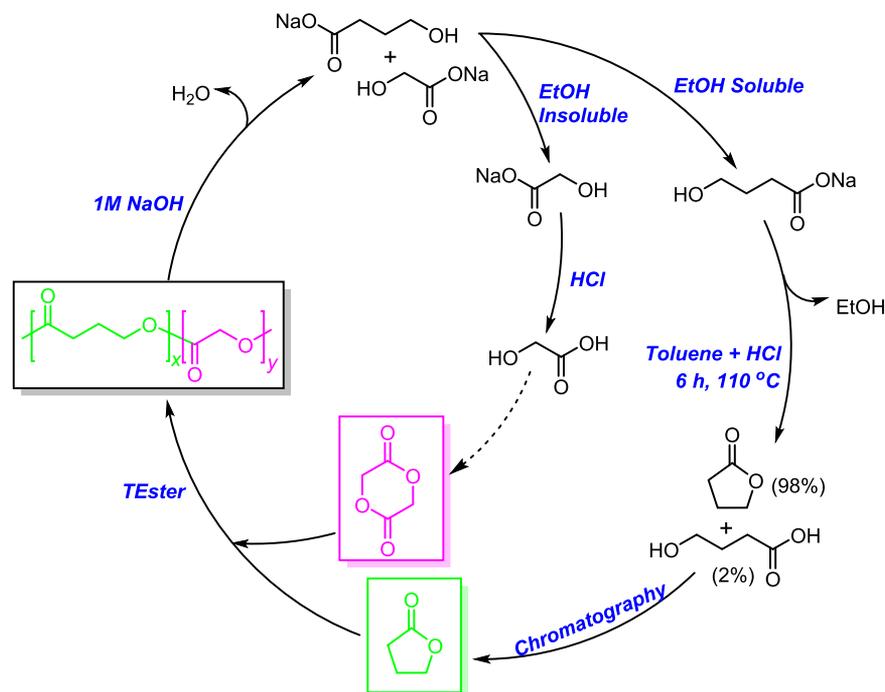
Scheme 2. Proposed Transition States and Calculated ΔG^\ddagger (kcal/mol) Values of Degradation Pathways of the Chains with Continuous G or B and Isolated G Sequences Calculated as Energy Differences between the Transition State and the Corresponding Relaxed Polymer Chain



facile, until the process leads to a more stable, terminal G-OH chain. Overall, the DFT results showed that the presence of the isolated G units (i.e., alternating G/B sequences) along the chain and, most relevantly towards the end, or preferably placed at the chain end, renders a more difficult decomposition process, thus enhancing the thermal properties of the copolymer possessing such sequences. Worth noting here is that the TS for intrachain backbiting, modeled by $\cdots\text{G-G-OR}$ and $\cdots\text{B-B-OR}$ (R = Me), has much higher ΔG^\ddagger energies, about 95 kcal/mol. This result agrees with the considerably higher thermostability observed for the cyclic PBL relative to its linear counterpart.¹⁵

To experimentally probe the influence of the isolated G sequence on the polymer thermal stability, we introduced a G unit into the PBL-OH chain end, namely, PBL-COCH₂OH, via the reaction of PBL-OH with 2-[(*tert*-butyldimethylsilyl)-oxy]acetyl chloride and subsequent deprotection. The successful chain-end modification was clearly shown by ¹H NMR analysis in combination with matrix-assisted laser desorption ionization time-of-flight mass spectrometry analysis (Figures S9 and S10). Importantly, the resultant PBL-COCH₂OH showed remarkably higher thermal stability (T_d = 271 °C) than PBL-OH (T_d = 208 °C) (Figure S11), thereby

Scheme 3. Chemical Recycling Route of poly(GA-co-BL)



further confirming that the isolated G sequence is responsible for enhancing the thermal stability.

Chemical Recyclability of Unconverted Poly(GA-co-BL). The chemical recyclability of poly(GA-co-BL) was examined by both thermolysis at high temperature (~ 300 °C, 12 h) and chemolysis in the presence of a simple metal salt [e.g., $\text{Bi}(\text{OTf})_3$, ZnCl_2] at milder temperature (120 °C, 12 h). However, neither of these approaches recycled the feedstock of the G units present in the copolymer, although B units in the copolymer can be quantitatively recovered in the form of a BL monomer (Figure S12). The desired glycolide was formed during thermolysis but was accompanied by glycolic acid and PGA oligomer, while during the chemolysis, only the formation of PGA oligomer was observed (Figure S12). To reduce the energy input and establish complete feedstock recycling, we then examined the hydrolysis pathway. Monitoring of the hydrolysis process in situ by ^1H NMR indicated that hydrolysis in acidic aqueous solution ($[\text{H}^+]/\text{D}_2\text{O} = 1.0$ M) led to only a small degree of degradation (Figures S13 and S17) and a trace of degradation products was detectable under the neutral condition (Figures S14 and S17) even after 16 days. In sharp contrast, basic aqueous conditions ($[\text{OH}^-]/\text{D}_2\text{O} = 1.0$ M) significantly enhanced the rate of hydrolytic degradation; within 2 h degradation of the copolymer was completed (Figure S16), affording sodium glycolate and sodium γ -hydroxybutyrate as clean products (Figures S15 and S17). Based on this finding, we developed a simple and efficient recycling route to recycle the feedstocks of poly(GA-co-BL), as depicted in Scheme 3. The resulting sodium glycolate and sodium γ -hydroxybutyrate can be easily separated due to large solubility difference in ethanol, where the ethanol-soluble fraction gave sodium γ -hydroxybutyrate, while the insoluble fraction gave sodium glycolate (Figure S18). After acidification, glycolic acid, the industrial raw material of glycolide and PGA, was obtained in quantitative yield (98%, Figure 3, top; Figure S19). The quantitative recovery of BL (98%, Figure S20) was achieved via cyclization

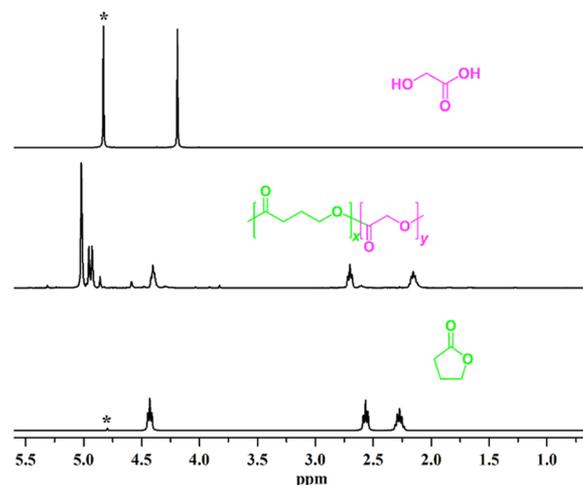


Figure 3. Overlays of ^1H NMR spectra: recovered glycolic acid (top, D_2O), starting poly(GA-co-BL) (middle, CF_3COOD), and recovered BL (bottom, D_2O). *The residual NMR solvent peak.

of sodium γ -hydroxybutyrate in toluene at 110 °C for 6 h under acidic conditions, bypassing γ -hydroxybutyric acid, which is a FDA-regulated substance and imposes potential dangers to human health if misused. After flash chromatography, pure BL was obtained in high isolated yield of 94% (Figure 3, bottom).

CONCLUSIONS

This work established the TEster as a convenient and effective method for achieving closed-loop upcycling of PGA in the industrial solvent BL via the formation of the copolymer bearing the property-enhancing isolated G sequences and full chemical recyclability. The successful example presented here has not only addressed the challenge encountered in the synthesis of poly(GA-co-BL) via coordination or anionic copolymerization of GA and BL due to their very large

reactivity difference but, more importantly, also led to the discovery of the isolated G sequence generated in the copolymer unique to the TEster approach, which is responsible for the enhanced thermal stability of the resulting copolymer. This upcycled copolymer can be fully recycled back to pure glycolic acid and BL building blocks using a simple and efficient recycling route, thus successfully constructing a closed-loop polymer upcycling scheme. Noteworthy also is the fact that this upcycling can also be performed using a mixture of homopolymers (Figures S21–S24). Future work will examine if the TEster strategy disclosed herein could also be applicable to other polyester/monomer pairs and homopolymer pairs and if advanced catalysts could be designed to enhance the TEster selectivity for generating property-enhancing sequences.

■ ASSOCIATED CONTENT

📄 Supporting Information

The Supporting Information is available free of charge on the ACS Publications website at DOI: 10.1021/acs.macromol.9b00817.

Materials, reagents, and methods; synthesis of PGA homopolymer; general transesterification procedures; polymer characterizations; chain-end modification of PBL; feedstock recycling of poly(GA-co-BL), computational details (PDF)

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Notes

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